



Chemical Changes

Specification statement	Self-assessment		
	First review 4-7 months before exam	Second review 1-2 months before exam	Final review Week before exam
These are the bits the exam board wants you to know, make sure you can do all of these...			
I can describe the reaction between metal and oxygen	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can recall the order of the reactivity series	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe when a displacement reaction might take place	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can use experimental data to work out the order of reactivity	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe how unreactive metals are found in the Earth	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe reduction	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe the process of extracting aluminium by electrolysis	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe oxidation as the loss of electrons Higher tier only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe reduction as a gain of electrons Higher tier only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can write balanced ionic half equations Higher tier only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can determine which element in a reaction is oxidised or reduced from the equation Higher tier only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can use the general equation to give the products from a reaction	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can determine the formula of a salt from common ions	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe how to make a pure salt	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can describe the ions that lead to acidic and alkaline conditions	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can use the pH scale to describe how acidic or alkaline a solution is	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can use an equation to show neutralisation	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can carry out a titration	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can calculate a concentration from titration data Chemistry only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹
I can give examples of strong and weak acids Higher tier only	☺ ☹ ☹	☺ ☹ ☹	☺ ☹ ☹



I can describe how concentration relates to pH Higher tier only	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can use the terms strong, weak, concentrated and dilute in term of acids Higher tier only	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can explain why compounds need to be molten or dissolved to conduct	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can describe the movement of ions during electrolysis	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can predict the products of electrolysis	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can write balanced half equations to describe what happens at each electrode	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can describe how to test for the production of chlorine gas	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can describe how to test for the production of hydrogen gas	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can describe how to test for the production of oxygen gas	😊 😐 😞	😊 😐 😞	😊 😐 😞
I can describe what happens to aqueous solutions that are electrolysed	😊 😐 😞	😊 😐 😞	😊 😐 😞